

UNDERSTANDING CALCIUM CARBONATE CRYSTALLIZATION PATHWAYS FOR CONCRETE RECYCLING

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The use of concrete has been growing continuously in the 20th and 21st century¹, leading to an increase of its environmental footprint. The current global warming crisis and the significant consumption of natural reserves have pushed the building industry to search for new and sustainable solutions. New breakthrough technologies for lowering the CO₂ footprint include a circular utilization of demolished concrete. Recent developments have shown that most of the CO₂ originally released by limestone calcination during clinker production can be sequestered by carbonation of the recycled cement paste^{2,3}. This technology, currently under development by the cement industry, consists in the carbonation of a part of recycled concrete in a humid/aqueous medium, leading to the formation of carbonate minerals and therefore to the permanent storage of the CO₂ in solid form.

A successful deployment of this technology at large scale needs a good understanding and, eventually, a control of the CaCO₃ polymorphism. Here, we will show a characterization of the carbonation reactions using both laboratory and synchrotron X-ray scattering experiments. The effect that different ions from cement hydrates –e.g., silicate, aluminate, magnesium- and different organics used as modifiers of the hydration kinetics – e.g., gluconate- have on the calcium carbonate crystallization kinetics will be presented and discussed in the framework of existing theories of mineral formation.

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