Reactive transport of Selenite and Strontium through a goethite coated sand column

LUKAS ZUNFTMEISTER¹, ZHE NIE², FRANK HEBERLING³, JOHANNES LUETZENKIRCHEN¹, NICOLAS FINCK³, CHUNLI LIU², REMI MARSAC⁴, NORBERT JORDAN⁵ AND KHALIL HANNA⁶

¹KIT

²Beijing National Laboratory for Molecular Sciences, Fundamental Science Laboratory, on Radiochemistry and

Radiation Chemistry, College of Chemistry and Molecular

Engineering, Peking University, China

³Karlsruhe Institute of Technology (KIT)

⁴Géosciences Rennes - UMR CNRS 6118

⁵Helmholtz-Zentrum Dresden-Rossendorf e.V. VAT No. DE 140 213 784

⁶Institut des Sciences Chimiques de Rennes - UMR CNRS 6226

Presenting Author: lukas.zunftmeister@kit.edu

A sound understanding of the key mechanisms governing the mobility of contaminants in the environment is of primary importance. Goethite occurs in numerous natural settings as well as in potential host rocks for high level nuclear waste repositories as an accessory mineral.

A previously proposed CD-MUSIC type model for Se(IV) adsorption on goethite in the presence of Sr(II) [1] was found to be insufficient for simulating breakthrough curves (BTCs) in column experiments, in which Se(IV) and Sr(II) containing solutions were pumped through a goethite-coated sand packed column. The simulations of the BTCs improved when contributions from the quartz surface were included (i.e. assuming that goethite does not completely cover the quartz surface). The BET surface area (SSA) is corrected, based on a new approach [2], where the crystal habit of a model crystal, derived from TEM images, is used. Simulation of the static data with this SSA and associated face contributions and site densities, starting from proton related surface charge and subsequently Se(IV) and Sr adsorption to goethite, yielded a new set of parameters for the goethite model. These parameters and inclusion of a silica acid-base and strontium adsorption model resulted in improved simulation results. The partial coverage of the sand with goethite poses further challenges due to charge heterogeneities and overlapping of surface potentials. Also, the dissolution of quartz and re-adsorption of silicic acid may play a significant role [3], which will be implemented in the model.

We acknowledge funding by the German federal ministry for education and research through the collaborative project KRIMI, grant agreement 02NUK056A.

References

- 1. Nie, Z., et al., Adsorption of Selenium and Strontium on Goethite: EXAFS Study and Surface Complexation Modeling of the Ternary Systems. Environmental Science & Technology, 2017. 51(7): p. 3751-3758.
- 2. Livi, K.J.T., et al., Crystal Face Distributions and

Surface Site Densities of Two Synthetic Goethites: Implications for Adsorption Capacities as a Function of Particle Size. Langmuir, 2017. 33(36): p. 8924-8932.

3. Jordan, N. et al., Competition between selenium (IV) and silicic acid on the hematite surface. Chemosphere 2009. 75(1): p. 129–134.