Hydroxyal Radicals produced from Oxygenation of Fe(II)-bearing Clay Minerals and the Resultant Oxidizing Impact

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Dark production of reactive oxygen species (ROS) has attracted great interests in recent years. Oxygenation of reduced component such as Fe(II) is an important source of ROS in nature. We recently reported the production of hydroxyl radicals (•OH) from oxygenation of sediment Fe(II) in subsurface when the redox condition was disturbed by O₂ [1]. Herein we show that Fe(II)-bearing clay minerals, i.e., reduced nontronite (NAu-2), can efficiently activate O₂ to produce •OH under neutral conditions. A stepwise oneelectron transfer process with the involvement of superoxide (O2) and hydrogen peroxide (H2O2) was identified for O2 reduction to •OH by the reduced NAu-2 (54.5% of Fe as Fe(II)). The electron transfer from structural Fe(II) in NAu-2 to O2 was explored through characterizing the reduced NAu-2 upon oxidation for different time by fourier transform infrared spectroscopy (FT-IR), Mössbauer spectra, X-ray photoelectron spectroscopy (XPS) and ultraviolet-visible spectrum (UV). Different coordination of structural Fe(II) showed different reactivities upon oxygenation, and both structrual Fe(II) and Fe(III) rearranged during the course of oxygeantion. We finally found that the •OH produced from oxygenation of chemically and biologically reduced NAu-2 led to the oxidative transformation of environmental contaminants like trichloroethylene, and from oxygenation of field sediment induced benzoic oxidation. Thus, we suggest that •OH produced from oxygeanting Fe(II)-bearing clay minerals could be an important mechanism for contamiannt attenuation in redox-dynamic conditions.

[1] Tong et al. (2016) Environ. Sci. Technol. 50, 214-221.